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Chapter 122 Chemical Calculations

In chemical calculations, mole ratios are used to convert between moles of reactant and moles of product, between moles of reactants, or between moles of products In a typical

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stoichiometric problem, the given quantity is first converted to moles.

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Terms in this set (12)

AT. In mass- mass calculations, the molar mass is used to convert mass to moles.

NT. The mole ratio 2 mol HF/1 mol SnF₂ can be used to determine the mass of SnF₂ produced according to the equation: Sn(s) + 2HF ...

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Section 122

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Key 122.6, HCl = 36.5,
ClO₂ = 67.5, Cl₂ =
71.0, H₂O = 18.0

g/mole) The amount of
Cl₂ produced is 18.7 g.
Calculate the percent
yield of Cl₂. 5) Ans:

80.6 % 6. 28.50 g
sample of a compound
of carbon, sulfur,

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hydrogen, and oxygen is burned. 35.25 g CO₂ and 14.65 g SO₂ are produced. Analysis of hydrogen

122 Chemical Calculations Worksheet Answer Key

Quantitative calculations involving reactions in solution are carried out in the same manner as we discussed in Chapter 11. Instead of masses,

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however, we use volumes of solutions of known concentration to determine the number of moles of reactants. Whether we are dealing with volumes of solutions of reactants or masses of reactants, the coefficients in the balanced chemical equation tell us ...

Chapter 12.2: Stoichiometry of Reactions in

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Solution ...

Chemistry (12th Edition) answers to Chapter 12 - Stoichiometry - 12.2 Chemical Calculations - Sample Problem 12.3 - Page 391 12 including work step by step written by community members like you.
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**Chapter 12 -
Stoichiometry - 12.2
Chemical
Calculations ...**

Chapter 3

Stoichiometry:

Calculations with

Chemical Formulas and
Equations John D.

Bookstaver St. Charles
Community College

Cottleville, MO

Chemistry, The Central
Science, 11th edition

Theodore L. Brown, H.

Eugene LeMay, Jr., and

Bruce E. Bursten

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Stoichiometry: Calculations with Chemical Formulas and ...

In a chemical reaction, one or more reactants are transformed into products: reactants \rightarrow products. The purpose of a chemical equation is to express this relation in terms of the formulas of the actual reactants and products that define a particular chemical change. For

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example, the reaction of mercury with oxygen to produce mercuric oxide would be expressed by the equation

Chemical Equations and Calculations

Chapter 3: Calculations
with Chemical
Formulas and
Equations Molecular
Mass and Formula
Mass for molecular
compounds: the
molecular mass is the

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mass (in amu) of one molecule of the compound
molecular mass = \sum atomic masses of elements present
ex. P 20 5
molecular mass =
 $2(30.97 \text{ amu}) + 5(16.00 \text{ amu}) = 141.94 \text{ amu}$
for ionic compounds:

Chapter 3: Calculations with Chemical

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Chemical Calculations
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and Chemical

Equations. 10.1

Equation Stoichiometry

371 There is a shortcut

for this calculation. We

can collapse all five of

the conversion factors

above into one. The

reaction equation tells

us that there are six

moles of H_2O for each

mole of P_4O_{10}

Chapter 10 Chemical Calculations and equations

12.2 Chemical

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Calculations > Writing and Using Mole Ratios
The figure below shows another way to represent the steps for doing mole-mass and mass-mass and mass-mole stoichiometric calculations.

- For a mole-mass problem, the first conversion is skipped.
- For a mass-mole problem, the last conversion is skipped.

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12.2 Chemical Calculations > - Useful Advice

Concept Review with Key Terms. The subject of stoichiometry involves quantitative calculations based on chemical formulas and chemical equations..

3.1 Molecular Masses and Formula Masses—Molecular masses and formula masses are the masses, expressed in atomic mass units (u), of individual molecules

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and formula units.

They are calculated from the masses of the atoms represented in the ...

Stoichiometry: Chemical Calculations

from Chemical
Manufacturing
Facilities August 2007
Final Prepared for
Emission Inventory
Improvement Program
Prepared by Mitchell
Scientific, Inc.

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Westfield, NJ RTI
International Research
Triangle Park, NC
Volume II: Chapter 16

EIIP Vol. II: Chapter 16 Methods for Estimating Air ...

The chemical reaction between these substances produces aluminum oxide, water, nitrogen gas, and hydrogen chloride. Although the solid rocket boosters each have a significantly

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lower mass than the liquid oxygen and liquid hydrogen tanks, they provide over 80% of the lift needed to put the shuttle into orbit—all because of chemical reactions.

Introduction to Chemical Reactions and Equations ...

The chemical formula, P_4O_{10} , provides such a conversion factor. It shows that there are four atoms of

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phosphorus for each
molecule of P_4O_{10} . 4
atoms P 1 molecule or
4 atoms P 1 molecule
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Chemical Calculations
and Chemical Formulas

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