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**Modern Chemistry
Section 15 3 Review
Answers**

Chemistry - Mid Term
Exam Review Sheet #1

The midterm exam covers chapters 1 - 4 & 9 - 11. You should read through each chapter, look over old tests you still have. answer the following questions and do the calculations in order prepare yourself for the mid-term. 1. Define tbc

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Following terms and describe where each is located.

**Chemistry - Mid
Term Exam Review
Sheet #1**

Modern Chemistry
Chapter Review 5
Answers Modern
Chemistry 37 The
Periodic Law CHAPTER
5 REVIEW The Periodic
Law SECTION 3 SHORT
ANSWER Answer the
following questions in
the space provided. 1.

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When an electron
is added to a neutral
atom, energy is (a)
always absorbed.
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Chapter 5 Review
Answers

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CHAPTER 1 REVIEW
Matter and Change
MIXED REVIEW SHORT

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ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance.

homogeneous a. sugar
homogeneous d.

plastic wrap
homogeneous b. iron
filings heterogeneous

e. cement sidewalk
heterogeneous c.

granola bar 2. For each
type of investigation,

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select the most
appropriate ...

**1 Matter and Change
- HUBBARD'S
CHEMISTRY**

CHAPTER 7 REVIEW

Chemical Formulas and
Chemical Compounds

SECTION 1 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

c In a Stock system
name such as iron(III)
sulfate, the Roman
numeral tells us (a)

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Answers

how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on each Fe ion.

7 Chemical Formulas and Chemical Compounds

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 3 SHORT

ANSWER Answer the following questions in the space provided. 1.

a The notation for

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sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b ...

6 Chemical Bonding

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Answer Key Modern
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Solutions CHAPTER 12

REVIEW Solutions

Teacher Notes and

Answers Chapter 12

SECTION 1 SHORT

ANSWER 1. c 2. a 3. b

2. a. alcohol b. water c.

the gels 3. The mixture

is a colloid. The

properties are

consistent with those

reported in Table 3 on

page 404 of the text.

The particle size is

small, but not too

small, and the mixture

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CHAPTER 12 REVIEW
Solutions

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Answers

questions about the data. Then they will ...

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Lesson Plans &
Worksheets
Reviewed by ...**

CHAPTER 9 REVIEW
Stoichiometry MIXED
REVIEW SHORT

ANSWER Answer the following questions in the space provided. 1.

Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is

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the value of the
coefficient x in this
equation? 40.07 g/mol
b. What is the molar
mass of C₃H₄? 2 mol
O 2:1 mol H 2O c. What
is the mole ratio ...

mc06se cFMsr i-vi

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Review of Chemistry of Matter (Chapter 2, & 3 Test Review) ANSWERS
Directions: Review and check all your answer with this answer sheet to get ready for the test on Chemistry of Matter (Chapter 2 & 3).
Holt McDougal Modern Chemistry Chapter 10: States of Matter ...

Chemistry Chapter 10 Test Review Answers

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Modern Chemistry

Chapter 15 Test

Review 25 multiple
choice definition of self-
ionization $[H_3O^+]$

$[OH^-] = 1 \times 10^{-14}$

definition and formula
for calculating $pH =$

$-\log[H_3O^+]$ $pH > 7$ is a

base; $pH < 7$ is an acid;

$pH = 7$ is neutral pH

range is normally 0 to

14 calculate pH from

$[H_3O^+]$ calculate

$[H_3O^+]$ from pH

definitions of

indicators, transition ...

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Modern Chemistry
Chapter 15 Acid-
Base Titration & pH

CHAPTER 5 REVIEW

The Periodic Law

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1.

c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original

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design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

5 The Periodic Law

Color Printable Periodic Table - Pretty much everything you need that can fit on a page and still be readable. Color table with atomic numbers, element

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symbols, element names, atomic weights, periods, and groups. [2013 Edition] [2012 Edition] Black/white Printable Periodic Table - Black/white table with atomic numbers, element symbols, element names, atomic weights, periods.

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grow old to spend to go
to the book foundation
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